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e4 Mathematics for Physical Chemistry thesineandcosinetotheappropriatenumberofdigits. (31) 2n rad 360 = 0.54 rad sin(30.5) = 0.5075 sin(31.5) = 0.5225 sin(31) = 0.51 cos(30.5) = 0.86163 cos(31.5) = 0.85264 cos(31) = 0.86 15. Some elementary chemistry textbooks give the value of R, the ideal gas constant, as 0.0821 l atm K-1 mol-1. a.

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Write this function in terms of Cartesian coordinates. $\cos(\theta) = z \sqrt{z^2 + y^2 + x^2}$ Mathematics for Physical Chemistry b. The ψ_{211} function is given by $\psi_{211} = \sqrt{8} \pi u^0 033/2 12$.

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e46 Mathematics for Physical Chemistry where M is the molar mass and NAv is Avogadro's constant: $v_p = 2mkBT = 2(1.3806568 \times 10^{-23} \text{ J K}^{-1})(298.15 \text{ K})(4.652 \times 10^{-26} \text{ kg})^{1/2} = 420.7 \text{ m s}^{-1}$ For a rough graph, we omit a constant factor and plot the function $g = u^2 e^{-u^2}$ where $u = v/v_p$. 31.

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Example: For example suppose we have $x = 3$, $y = 5$ and $a = 6$ then: 1. $a(x + y) = 6(3 + 5) = 6 \cdot 8 = 48$ 2. $(a \cdot x) + (a \cdot y) = (6 \cdot 3) + (6 \cdot 5) = 18 + 30 = 48$ So $a(x + y) = (a \cdot x) + (a \cdot y)$ and the distributive rule holds. Factorisation. Factorisation is used to tidy up equations in order to make them easier to read and understand.

Maths for Chemists

Solutions to exercises 1B.1(b) The mean speed is $8RT/v_{mean}$ The mean translational kinetic energy is $m \cdot 3RT/2$ The ratios of species 1 to species 2 at the same temperature are $E_k/v_{mean,1} \cdot v_{mean,2} = (1/2) m_1 v_{mean,H}^2 / m_2 v_{mean,Hg}^2 = 200.6 \cdot 4.003 / 2 \cdot 200.6 = 1.0015$ (ii) The mean translation kinetic energy is independent of molecular mass and depends upon temperature alone!

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