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## Guide

[FREE] Unit 3 Chemical Bonding Study Guide Answer Key. Chemical Bonds Study Guide Define the following:

1. Chemical formula - the way of expressing information about the proportions of atoms that constitute a particular chemical compound, using element symbols and numbers.

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### Answers - 09/2020

From a general summary to chapter summaries to explanations of famous quotes, the SparkNotes Introduction to Chemical Bonding Study Guide has

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## **Introduction to Chemical Bonding: Study Guide | SparkNotes**

Characteristics of both bonds:

- \* Occur between 2 atoms
- \* Composed of 2 electrons
- \* Have both ionic and covalent characteristics
- \* Together = 100%
- \* Both bonds are measured on an electronegativity scale
- \* Both contain a nonmetal
- \* Chemical bonds are determined by using the "magic number" (1.67)
- \* Have

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bond angle and bond axis

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Metals (cations)

Nonmetals (anions) Cations An atoms that has lost or gained electrons creating a charge. Loses electrons

Anions Gains electrons

Charge that ions normally make when bonding Most atoms need 8 valence electrons in their outermost shells in order to be stable. To show the charge of multivalent ions.

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A. Fill in the Blank 1. The

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electrons involved in the formation of a chemical bond are called 2. A chemical bond that results from the electrostatic attraction between positive and bond negative ions is called a(n) 3. If electrons involved in bonding spend most of the time closer to one atom rather than

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1. for most combinations of atoms a chemical bond will occur if the potential energy is reduced by the change a. most atoms have lower potential energy in bonds 2. chemical bonding is often an energy releasing process 3. breaking of



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chemical bonds is often an energy absorbing process

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When an atom interacts with another atom, it forms what is called a chemical bond. This interaction, this chemical bond links the two atoms together into something called a molecule. Making a...

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What is a chemical bond? a mutual electrical attraction between the nuclei and valence  $e^-$  of different atoms that binds the atoms

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together What happens to the valence electrons in an ionic bond vs. a covalent bond? In an ionic bond, they are transferred.

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Chemical Bonding n lesson 1 you have learnt about the structure of atom while in the lesson 2 you studied about the classification of elements and the variation in atomic properties. You know that molecules are obtained by the combination of two or more than two atoms of the same or different elements. In this lesson you will study

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...

Covalent bonding represents a situation of about equal sharing of the electrons between nuclei in the bond. Covalent bonds are formed between atoms of approximately equal electronegativity. Because each atom has near equal pull for the electrons in the bond, the electrons are not completely transferred

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from one atom to another.

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Science (112): Test Practice  
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Chemical Bonds Study Guide  
11 November 2016 A molecule with covalent bond o Formula unit with ionic bond \*  
Molecules: formed by covalent bonds \* Lattice energy is the energy released in the formation of an ionic compound.

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